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$\Delta T_b = K_b m$ . where  $m$  is the molality of the solution and represent moles of solute in 1 kg of solvent and  $K_b$  is called molal boiling point elevation constant or molal boiling point constant or ebullioscopic constant.. If  $m=1$ . then.  $\Delta T_b = K_b$ . Thus, molal boiling point elevation constant,  $K_b$  is defined as the elevation in boiling point for 1 molal solution i.e. a solution containing 1 ...

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