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occupies a volume of
458 mL at a pressure of

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LOI kPa and
temperature of 295 K.
When the pressure is
changed, the volume
becomes 477 mL. If
there has been no
change in temperature,
what is the new
pressure? A gas
occupies a volume of
2.45 L at a pressure of
1.03 atm and a
temperature of 293 K.

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the factory is 761.3 mm Hg. The pressure inside the bag is the same.

What is the pressure inside the bag of potato Chips in Pa? The same bag of potato chips from Problem 5 is shipped to a town in Colorado, where the atmospheric pressure is 99.82 kPa. What is the difference (in Pa) between the

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pressure in the bag and
the atmospheric

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A gas occupies a volume of 458 mL at a pressure of 101 kPa and temperature of 295 K.

When the pressure is changed, the volume becomes 477 mL. If there has been no change in temperature, what is the new pressure?

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temperature of 295-K.
When the pressure is
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becomes 477-mL. If
there has been no
change in temperature,

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what is the new
pressure? 6. What will
be the volume of a gas
sample at 309-K if its
volume at 215-K is
3.42-L?

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13.3: Factors Affecting
Solubility Conceptual

Problems Use the
kinetic molecular theory
of gases discussed in
Chapter 10 to explain
why the solubility of
virtually all gases in
liquids decreases with
increasing temperature.

13.E: Properties of

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1. What volume would be occupied by 100. g of oxygen gas at a pressure of 1.50 atm and a temperature of 250°C? $n = \frac{m}{M} = \frac{100. \text{ g}}{32.0 \text{ g/mol}} = 3.125 \text{ mol}$
(Use $T = 273 \text{ K}$)

2. An air-filled balloon has a volume of 5 L at 0.94 atm and 250°C. Soon after,

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the pressure changes to
0.99 and the
temperature changes to
OCC. What is the new
volume of the balloon?
.26 1.26 L ...

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empirical formula of a
compound, given that
the compound is found
to be 47.9 percent zinc
and 52.1 percent
chlorine by mass. 20.

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