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$[\text{Ni(s)} \rightarrow \text{Ni}^{2+} + 2\text{e}^{-}] \times 5$; red: $[\text{5e}^{-} + 8\text{H}^{+} + \text{MnO}_4^{-} \rightarrow \text{Mn}^{2+} + 4\text{H}_2\text{O}] \times 2$. balanced: $5\text{Ni(s)} +$

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balanced: $2\text{Ag(s)} + \text{Cu}^{2+}(\text{aq}) \rightarrow 2\text{Ag}^{+}(\text{aq}) + \text{Cu(s)}$ $E^{\circ} = E^{\circ}_{\text{ox}}(\text{Ag}) + E^{\circ}_{\text{red}}(\text{Cu}^{2+}) = 0.80\text{V} + 0.34\text{V} = 0.46\text{V}$. Not spontaneous. ox: $[\text{Ni(s)} \rightarrow \text{Ni}^{2+} + 2\text{e}^{-}] \times 5$; red: $[\text{5e}^{-} + 8\text{H}^{+} + \text{MnO}_4^{-} \rightarrow \text{Mn}^{2+} + 4\text{H}_2\text{O}] \times 2$. balanced: $5\text{Ni(s)} + 16\text{H}^{+}(\text{aq}) + 2\text{MnO}_4^{-}(\text{aq}) \rightarrow 5\text{Ni}^{2+}(\text{aq}) + 2\text{Mn}^{2+}(\text{aq}) + 8\text{H}_2\text{O}$.

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second force,

~~Chapter 18 Nuclear Chemistry~~

a) the energy given off when reactants collide b) *an energy barrier between reactants and products the heat released in a reaction* c) generally very high for a reaction that takes place rapidly

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