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ox: [Ag + + e - ]  $\times$  2; red: Cu2 + + 2e - Cu. balanced: 2Ag(s) + Cu2 + (aq) 2Ag + (aq) + Cu(s) E = E ox(Ag) + E red(Cu2 +) = - 0.80V + 0.34V = - 0.46V. Not spontaneous. ox: [Ni(s) Ni2 + + 2e - ]  $\times$  5; red: [5e - + 8H + + MnO - 4 Mn2 + + 4H 2O]  $\times$  2.

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Modern Chemistry 149 Chemical Equilibrium CHAPTER 18 REVIEW Chemical Equilibrium SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. Match the solution type on the right to the corresponding relationship between the ion product and the Ksp for that solution, listed on the left.

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Chapter 18 – Nuclear Chemistry231. 62. Describe the two opposing forces between particles in the nucleus, and with reference to these forces, explain why the ratio of neutrons to protons required for a stable nuclide increases as the number of protons in a nucleus increases.

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