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Formulas. STUDY. PLAY. Monatomic
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Ion. Consists of a single atom with a positive or negative charge resulting from the loss of gain of one or more valence electrons and behaves as a unit, yet it still has a charge. (+ or -). Determined by the amount of electrons lost or gained.

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or gain of electrons, respectively.

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for an acid normally starts with and H. When naming acids, you should first determine whether or not the acid contains oxygen.

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Chapter 9 Chemistry Chemical Names and Formulas, Acids, base, Cation, Ionic compounds. Compounds that contain one or more hydrogen atoms and produce.... An ionic compound that produces hydroxide ions when dissolved.... Any atom or group of atoms that has a positive charge.

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Compounds composed of cations and anions.

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Chapter 9 Chemical Names and
Formulas83 SECTION 9.3 NAMING
AND WRITING FORMULAS FOR
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MOLECULAR COMPOUNDS (pages 268-270) This section explains the rules for naming and writing formulas for binary molecular compounds. Naming Binary Molecular Compounds (pages 268-269) 1. Circle the letter of the type(s) of elements that form binary molecular compounds.

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Name Date Class CHEMICAL NAMES AND FORMULAS 9 SECTION 9.2 NAMING AND WRITING FORMULAS FOR IONIC COMPOUNDS I. Write the formulas for these binary ionic compounds. c. potassium iodide 2. write the formulas Page 25/38

for the compounds formed c. b. f. 3. Name the following inary ionic compounds. sodium sulfide g. CuCl: h. snc.l. a. Mn02 c. d. SrBr2 221 Chapter 9 Chemical Names and Formulas

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Chapter 9 Chemical Names and Formulas 9.1 Naming lons 9.2 Naming and Writing Formulas for Ionic Compounds 9.3 Naming and Writing Formulas for Molecular Compounds 9.4 Naming and Writing Formulas for Acids and Bases 9.5 The Laws Governing How Compounds Form Page 27/38

Chemical Names and Formulas Pittsfield Chapter 9 "Chemical Names and Formulas" ... in parenthesis after the name of the metal (Table 9.2, p.255) Predicting Ionic Charges Some of the post-transition elements also have Page 28/38

more than one possible oxidation state. Tin (II) = Sn2+ Lead (II) = Pb2+ Tin (IV) = Sn4+ Lead (IV) = Pb 4+

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a hydroxide ion b. lead(IV) ion c. sulfate ion d. oxide ion 47. zero; A compound is electrically neutral. 48. The symbols for the cation and

CHAPTER 9 Study Guide Quia Chapter 9 Practice Problems: Chemical Names and Formulas.

Section 9.1: Naming lons. 1. What is the charge on the ion typically formed by each element? a. oxygen c. sodium e. nickel, two electrons lost . b. iodine d. aluminum f. magnesium. 2. How many electrons does the neutral atom gain or lose when each ion forms? a. Cr3+ c. Li+ e. Cl- b. P3 ... Page 33/38

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Formulas" Tools. Copy this to my account; E-mail to a friend; Find other activities; ... In samples of any chemical compound, the masses of the elements are always in the same proportions any atom or group of atoms that has a negative charge ... place cation name first followed by the Page 36/38

anion name: naming polyatomic ...

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