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Monatomic Ion. Consists of a single atom with a positive or negative charge resulting from the loss of gain of one or more valence electrons and behaves as a unit, yet it still has a charge. (+ or -). Determined by the amount of electrons lost or gained.

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single atom with a positive
or negative charge resulting
from the loss or gain of
electrons, respectively.

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Writing Formulas for Acids
and Bases. An acid is a
compound that produces H^+
ions when it dissolves in
water. The formula for an
acid normally starts with

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Formulas. When naming acids, you should first determine whether or not the acid contains oxygen.

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Names and Formulas. Acids.
base. Cation. Ionic
compounds. Compounds that
contain one or more hydrogen

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An ionic compound that produces hydroxide ions when dissolved... Any atom or group of atoms that has a positive charge. Compounds composed of cations and anions.

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83 SECTION 9.3
NAMING AND WRITING FORMULAS FOR MOLECULAR COMPOUNDS
(pages 268–270) This section explains the rules for naming and writing formulas for binary molecular compounds. Naming Binary Molecular Compounds (pages

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268–269) 1. Circle the letter of the type(s) of elements that form binary molecular compounds.

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NAMES AND FORMULAS 9
SECTION 9.2 NAMING AND
WRITING FORMULAS FOR IONIC~~

COMPOUNDS I. Write the formulas for these binary ionic compounds. c. potassium iodide 2. write the formulas for the compounds formed c. b. f. 3.

Name the following binary ionic compounds. sodium sulfide g. CuCl_2 : h. SnCl_4 . a. MnO_2 c. d. SrBr_2 221

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Naming and Writing Formulas
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Naming and Writing Formulas
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Naming and Writing Formulas
for Acids and Bases 9.5 The
Laws Governing How Compounds
Form

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and Formulas” ... in
parenthesis after the name
of the metal (Table 9.2,
p.255) Predicting Ionic
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Transition elements also have more than one possible oxidation state. Tin (II) = Sn^{2+} Lead (II) = Pb^{2+} Tin (IV) = Sn^{4+} Lead (IV) = Pb^{4+}

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281 CHAPTER 9 Assessment 42.
a. 2- b. 1+ c. 1- d. 3+ 43.
a. 2+ b. 2+ c. 3+ d. 1+ 44.
a. barium ion b. iodide ion
c. silver ion d. mercury(II)
ion 45. cyanide, CN^- and
hydroxide, OH^- 46. a.
hydroxide ion b. lead(IV)
ion c. sulfate ion d. oxide
ion 47. zero; A compound is
electrically neutral. 48.
The symbols for the cation
and

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Chapter 9 Practice Problems:
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Section 9.1: Naming Ions. 1.~~

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What is the charge on the ion typically formed by each element? a. oxygen c. sodium e. nickel, two electrons lost . b. iodine d. aluminum f. magnesium. 2. How many electrons does the neutral atom gain or lose when each ion forms? a. Cr^{3+} c. Li^+ e. Cl^- b. P^3 ...

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