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Both:  $\text{Sn}^{2+} + (\text{E}_{\text{red}} = -0.14\text{V}; \text{E}_{\text{ox}} = -0.15\text{V})$ ,  $\text{Fe}^{2+} + (\text{E}_{\text{red}} = -0.44\text{V}, \text{E}_{\text{ox}} = -0.77\text{V})$ ,  $\text{Cl}_2 (\text{E}_{\text{red}} = 1.36\text{V}, \text{E}_{\text{ox}} = -1.47\text{V})$ . Oxidizing agents (list in order of decreasing  $\text{E}_{\text{red}}$ ):

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$\text{Cl}_2 > \text{H}^+ > \text{Sn}^{2+} >$

$\text{Fe}^{2+}$ . Reducing agents  
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$\text{Mg} > \text{Ni} > \text{Sn}^{2+} > \text{Fe}^{2+} > \text{Cl}_2 > \text{Au}$ . 4.

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