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 $2 \text{SO}_2 + \text{O}_2 \rightarrow 2 \text{SO}_3$ $\text{SF}_4 + 3 \text{H}_2\text{O} \rightarrow \text{H}_2\text{SO}_3 + 4 \text{HF}$
 $2 \text{KNO}_3 + 10 \text{K} \rightarrow 5 \text{K}_2\text{O} + \text{N}_2$.

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ANSWER: Correct Chapter 9 Reading Question 2

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Part A Iron reacts with oxygen in the air to produce rust (Fe_2O_3) according to the reaction shown below. $4 \text{ Fe} + 3 \text{ O}_2 \rightarrow 2 \text{ Fe}_2\text{O}_3$ How many moles of rust can be formed if 1.8 mols of iron are present? Hint 1. Begin with your given number of moles.

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Answers To Mastering Chemistry Chapter 8 CHAPTER 8 REVIEW Chemical Equations and Reactions SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. List four metals that will not replace hydrogen in an acid. Choose from Cu, Ag, Au, Pt, Sb, Bi, and Hg.

Mastering Chemistry Chapter 8 Answer Key

Mastering Chemistry Chapter 2: Rank from most to fewest valence electrons of the following elements: Kr, Ga, Br, As, K. - Answer: Kr (8), Br (7), As (5), Ga (3), K (1). Valence electrons are equal to the atom's main group

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