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The answer lies in the fact that some of the energy of the excited molecule is converted to a different (non-photon or lower energy photon) form. Some energy can go to a visible photon while another portion of the energy goes to a neighboring atom or molecule through atomic collisions.

Title: The Phantastic Photon and LEDs (combined lab)

Page 1 Answers \u2022 Phantastic Photon and LEDs. Phantastic Photon and LEDs. Answers to Questions: Phantastic Photons A. Glow-in-the-right color. Note #1: the LEDs are not monochromatic; the listed wavelengths are the peaks of the intensity spectrums provided by the manufacturer. Note #2: Although not the point of this lab, it is worth pointing out to students that we are using the first order antinodes, but higher orders are clearly visible.

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10) A typical IR wavelength could be from just under a micrometer to a milimeter. Use 750nm to get a photon that is almost visible. Calculate its frequency using (freq)(wavelength) = c. Then...

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