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Specific Heat Worksheet 2 Answers

Specific Heat Worksheet #2 Name: \_\_\_\_\_ Per: \_\_\_\_ Seat: \_\_\_\_

Directions: Calculate the following showing ALL work to receive credit. Formula  $Q = mc \Delta T$ , where Q is heat in joules, c is specific heat capacity in J/g C, m is the mass in grams, and delta T is the change in temperature in C. Q Work Answer with Units! 1 How much heat is lost when a 640 g piece of copper cools from 375 ° C, to 26 ° C?

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Specific Heat Worksheet #2.pdf - Specific Heat Worksheet#2 ...

Name Answer Key Date 9/9/15 Chp 2-1: Specific Heat Worksheet

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## Answers

(m) (  $\Delta T$  ) (C sp )=Q

1. Specific heat is the amount of energy that it takes to raise the temperature of 1 gram of a substance by 1 degree kelvin
2. Absolute zero is the temperature at which all molecular motion ceases
3. Endothermic process is a change in matter in which energy is absorbed
- 4.

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Specific Heat WS Answers - Name Answer Key Date Chp 2-1 ...

Chapter 10 Worksheet #2

1. Calculate the energy require (in calories) to heat 10.4 g of mercury from 37.0 oC to 42.0 oC.

Specific heat of mercury is 0.14 J/g oC.  $q = m c \Delta t$   
 $q = 10.4 \text{ g} \cdot 0.14 \text{ J/g oC} \cdot 5.00 \text{ oC} = 7.28 \text{ J} \cdot 1 \text{ cal} = 1.74 \text{ cal}$   
4.184 J

2. If 50. J of heat are applied to 10. g of iron, by how much will the temperature of the iron

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## Chapter 10 Worksheet #2 Answer

(ANSWERS) 1. A 500 g piece of iron changes  $7^{\circ}\text{C}$  when heat is added. How much heat energy produced this change in temperature? (Ans. 2,000 J) 2. When 300. cal of energy is lost from a 125 g object, the temperature decreases from  $45.0^{\circ}\text{C}$  to  $40.0^{\circ}\text{C}$ . What is the specific heat of this object? (Ans.  $0.48\text{ cal/g }^{\circ}\text{C}$  or  $2.0\text{ J/g }^{\circ}\text{C}$ )

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## Honors Chemistry Worksheet – Specific Heat

Before discussing Calculating Specific Heat Worksheet Answers, you need to recognize that Knowledge can be your answer to a

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better the next day, along with studying doesn't just stop the moment the school bell rings. Of which getting claimed, many of us provide you with a a number of basic yet helpful posts along with design templates made ideal for almost any educative purpose.

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Calculating Specific Heat Worksheet Answers | [akademiexcel.com](http://akademiexcel.com)

$q$  = amount of heat (J)  $m$  = mass (grams)  $c$  = specific heat (J/g ° C)

$T$  = change in temperature ( ° C) 2. Heat is not the same as temperature, yet they are related. Explain how they differ from each other. Heat is a combination of kinetic energy (measured by temperature) and potential energy. a. Perform calculations using:  
( $q = m c \Delta T$ ) b.

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## Worksheet- Calculations involving Specific Heat

Answers are provided at the end of the worksheet without units. 1.

A 15.75-g piece of iron sorbs 1086.75 joules of heat energy, and its temperature changes from 25 to 175°C. Calculate the specific heat capacity of iron. = 'C ' Q 5) 2. How many joules of heat are needed to raise the temperature of 10.0 g of

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## Specific Heat Wksht20130116145212867

Two page worksheet using Specific Heat Capacity. Questions start easy then become gradually harder. Answers included on separate sheet. Also includes a spreadsheet to show how the calculations have been done.



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## Answers

with all the basic needs for an air conditioning system. When you are looking for a sheet, you can also check online to determine what works best for you.

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### Specific Heat Problems Worksheet Answers

Heat Transfer/ Specific Heat Problems Worksheet Solving For Heat (q) 1. How many joules of heat are required to raise the temperature of 550 g of water from 12.0 °C to 18.0 °C? 2. How much heat is lost when a 64 g piece of copper cools from 375 °C, to 26 °C? (The specific heat of copper is 0.38452 J/g x °C). Place your answer in kJ. 3.

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## Answers

### Heat Transfer/ Specific Heat Problems Worksheet

For the last step, with proper sig figs, I get 91.2, which is essentially the same answer as step #3. This is because of the small specific heat of the aluminum ( $0.089 \text{ J/g } ^\circ\text{C}$ ). 19. A sample of cobalt, A, with a mass of 5.00 g, is initially at  $25.0 \text{ } ^\circ\text{C}$ . When this sample gains 6.70 J of heat, the temperature rises to  $27.9 \text{ } ^\circ\text{C}$ .

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### Thermochem Worksheet #2 Answers - ChemTeam

What is the specific heat of an unknown substance if a 2.50 g sample releases 12 calories as its temperature changes from  $25 \text{ } ^\circ\text{C}$  to  $20 \text{ } ^\circ\text{C}$ ? ANSWER KEY. HEAT Practice Problems .  $Q = m \times c \times \Delta T$   
 $5.0 \text{ g of copper was heated from } 20 \text{ } ^\circ\text{C to } 80 \text{ } ^\circ\text{C}$ . How much energy was used to heat Cu? (Specific heat capacity of Cu is  $0.092$

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## Answers

cal/g ° C) 27.6 cal

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### HEAT Practice Problems

Worksheet- Calculations involving Specific Heat 1. For  $q = m c \Delta T$ : identify each variables by name & the units associated with it. 2. Heat is not the same as temperature, yet they are related. Explain how they differ from each other. (-m.c.AT) a. Perform calculations using 1. Gold has a specific heat of  $0.129 \text{ J/(g}\cdot\text{°C)}$ . How

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North St. Paul-Maplewood Oakdale / Overview

Specific Heat and Heat Capacity Worksheet DIRECTIONS: Use  $q = (m)(C_p)(\Delta T)$  to solve the following problems. Show all work and

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## Answers

units. Ex: How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22 ° C to 55 ° C, if the specific heat of aluminum is 0.90 J/g ° C? 1.

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Specific Heat and Heat Capacity Worksheet

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Campbell Science - Home

For  $q = m c \Delta T$  : identify each variables by name & the units associated with it.  $q$  = amount of heat (J)  $m$  = mass (grams)  $c$  = specific heat (J/g ° C)  $\Delta T$  = change in temperature ( ° C) 2. Heat is not the same as temperature, yet they are related. Explain how

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they differ from each other.

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## Chemistry Specific Heat Worksheet Answers

Here are the heat capacities of the four substances:  $0.10 \text{ cal/g } ^\circ \text{C}$ ,  $0.25 \text{ cal/g } ^\circ \text{C}$ ,  $1.0 \text{ cal/g } ^\circ \text{C}$ , &  $0.2 \text{ cal/g } ^\circ \text{C}$ . Match & then label each substance with its specific heat capacity on the graph. See graph above. 7. If something has a high specific heat capacity will it take a lot of heat or a little heat to change its temperature? Explain

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