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Stoichiometry Workbook Answers

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2. Stoichiometry is based on the law of conservation of mass. _____
3.

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Chapter 12 Stoichiometry127 SECTION 12.1 THE ARITHMETIC OF EQUATIONS (pages 353–358) This section explains how to calculate the amount of reactants required or product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

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Writing and Using Mole Ratios Using Visuals Figure 12.4
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Guided Reading And Study Workbook Chapter 38 Answers

Stoichiometry 353 12.1 FOCUS Objectives 12.1.1 Explain how balanced equations apply to both chemistry and everyday situations. 12.1.2 Interpret balanced chemical equations in terms of moles, representative particles, mass, and gas volume at STP. 12.1.3 Identify the quantities that are always conserved in chemical reactions.

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