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REACTIONS

Elementary Rate

Laws – Unimolecular,

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Reactions – Chemical

Kinetics Unimolecular

reaction-02 Collision

Theory Model, Rates

of Reaction,

Activation Energy,

Arrhenius Equation -

Chemical Kinetics

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Chemistry Zero Order

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Theory of

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Reactions...

A satisfactory theory

of these reactions

was proposed by F. A.

Lindemann in 1922.

According to

Lindemann, a

unimolecular

reaction  $A$

$\longrightarrow P$

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PDF Theory

proceeds via the  
following

mechanism:  $A + A$   
/rightleftharpoons

$A^* + A$  Here the

rate constants being

$k_f$  for forward

reaction &  $k_b$  for

backward reaction

and

The Lindemann

Theory of

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The Journal of

Physical Chemistry A

2007, 111 (19) ,

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Multichannel

Thermal

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Reactions. 2.

Application to the

Thermal Dissociation

of Formaldehyde. The

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Journal of Physical  
Chemistry A 2005,  
109 (37) , 8320-8328.

DOI:

10.1021/jp051027d.

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channel thermal  
unimolecular  
reactions. 1 ...

Description Theory of  
Unimolecular  
Reactions provides a  
comprehensive

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analysis of the theory of unimolecular reactions, also known to kineticists as the Rice-Marcus or the Rice-Ramsperger-Kassel-Marcus theory, and to those working in mass spectrometry and related fields as the quasi-equilibrium theory or the theory of mass spectra.

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1973) [WorldCat.org]

The master equation

of thermal

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reactions in the



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fall off range has been solved for a number of representative molecular systems. Weak collision broadening factors FWC ( $k_0 / k_{\infty}$ ) are derived and represented empirically. Weak collision efficiencies  $c$  for the low pressure range are

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calculated for very  
high temperatures.

Theory of Thermal

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Reactions in the

Fall off ...

Collision Efficiency

Parameter Influence

on Pressure-

Dependent Rate

Constant Calculations

Using the SS-QRRK

Theory. The Journal

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of Physical Chemistry

A 2020, 124 (31) ,

6277-6286. ...

Simplified Analysis

and Representation

of Multichannel

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Reactions. The

Journal of Physical

Chemistry A 2019,

123 (5) , 1007-1014.

Predictive

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possibilities of  
unimolecular rate  
theory | The ...

In chemical kinetics,  
the Lindemann  
mechanism is a  
schematic reaction  
mechanism for  
unimolecular  
reactions. Frederick  
Lindemann and J. A.  
Christiansen  
proposed the  
concept almost

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simultaneously in

1921, and Cyril

Hinshelwood

developed it to take

into account the

energy distributed

among vibrational

degrees of freedom

for some reaction

steps. It breaks down

an apparently

unimolecular

reaction into two

elementary steps,

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with a rate constant  
for each elementary  
step. The rate law  
and rate equation

Chemistry

Lindemann  
mechanism -

Wikipedia

Transition States in  
Unimolecular

Reactions, Theory of  
Unimolecular

Reactions, 10.1016/B  
978-0-12-262350-9.5

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0020-7,(344-393),  
(1973). Crossref  
Progress in Physical  
Organic Chemistry,  
Volume 9

## Wendell

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Reactions - Willcott -  
1972 ...

Until about 1921,  
chemists did not  
understand the role  
of collisions in

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### Unimolecular

processes. It turns out that the mechanisms of such reactions are actually quite complicated, and that at very low pressures they do follow second-order kinetics. Such reactions are more properly described as pseudounimolecular.

The cyclopropane



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### Unimolecular

described in Example 1 is typical of many decomposition reactions found to follow first-order kinetics, implying that the process is unimolecular.

#### 6.1.6: The Collision Theory - Chemistry LibreTexts

$$d[A] dt =$$

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$k_1[A][M]$ (forward first step) -  $k_{-1}[A$

$][M]$ (reverse first step) -  $k_2[A$

$]$ (forward second step) According to the steady-state approximation,  $d[A$

$]/dt = 0$ . Therefore the rate of production of A (first term in Equation 29.6.3) equals the rate of consumption

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(second and third terms in Equation 29.6.3 ):

29.6: The Lindemann Mechanism -

Chemistry LibreTexts

unimolecular

reactions in vapour

phase: 1)

LINDEMANN THEORY:

According to this

theory a

unimolecular

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## PDF Theory

reaction  $A \rightarrow P$

proceeds via  
following

mechanism:  $A^* \rightarrow P$

Product. Where,

$A^*$  = energized A

molecule that has

acquired sufficient

energy to enable it to

isomerizes or

decomposed i.e., the

vibration.

UNIMOLECULAR

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REACTIONS IN GAS

PHASE | Download

toppers notes

The Rice–Ramsperger

r–Kassel–Marcus

theory is a theory of

chemical reactivity. It

was developed by

Rice and Ramsperger

in 1927 and Kassel in

1928 and generalized

in 1952 by Marcus

who took the

transition state

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theory developed by Eyring in 1935 into account. These methods enable the computation of simple estimates of the unimolecular reaction rates from a few characteristics of the potential energy surface.

RRKM theory -

Wikipedia

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Unimolecular

Reactions of Peroxy  
Radicals in

Atmospheric

Chemistry and

Combustion. ...

MESMER: An Open-  
Source Master

Equation Solver for

Multi-Energy Well

Reactions, The

Journal of Physical

Chemistry A,

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116, 38, (9545-9560),  
(2012 ... Journal of  
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7, 5, (1244-1252),  
(2011 ...

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- Chemistry Europe  
The quantum  
tunneling  
contribution to the



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reaction rate constants is calculated using the semiclassical transition state

theory (SCTST)

developed by Miller and co-workers.

26–31 The SCTST calculates the overall quantum tunneling probability by summing over a series of TS

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vibrational configurat  
ion-dependent  
tunneling  
probabilities.

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