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There are many different  
ways of expressing the  
concentration of a given

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solution. Some of the most common include molarity, weight by volume, volume by volume and weight by weight. Weight by volume percent (w/v %) tells you the mass of solute in grams that has been added to a 100 mL

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A percent w/v solution is calculated with the following formula using the gram as the base measure of weight (w):  $\% w/v = g \text{ of}$



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solute/100 mL of solution.  
Example 1: Physiologic or isotonic saline is a 0.9% aqueous solution of NaCl. 0.9% saline = 0.9 g of NaCl diluted to 100 mL of deionized water,

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~~Calculating Percent  
Weight/Volume (% w/v)  
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For example, to find the %  
w/v of a solution the  
calculation is: (Mass of  
Solute (g) / Volume of  
Solution (ml)) x 100.

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Therefore, to figure out the % w/v of a 100ml solution that is made up of 65g nitric acid, we would divide 65g by 100ml and then multiply the answer by 100. This tells us that there is a nitric acid solution of

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65% w/v.

~~What Do % V/V, % W/W and %  
W/V Mean?~~

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NY Slip Op 50219 (U) [\*1]  
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on February 18, 2020 Supreme Court, Kings County Thomas, J. Published by New York State Law Reporting Bureau pursuant to Judiciary Law § 431.

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[\*1] Matter of Jefferson W.  
v Latonya P. 2018 NY Slip Op  
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Because percent solutions  
can be expressed in three

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different ways, it is imperative that the type of percent solution be explicitly stated. If this information is not provided, the end user is left to "guess" whether w/v %, w/w %, or v/v % was used. Each

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percent solution is  
appropriate for a number of  
different applications.

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challenges they face.

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Solutions that require  
materials from natural  
sources are often prepared  
w/v because the molecular

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formula of the substance is unknown and/or because the substance cannot be described by a single formula. A one percent solution is defined as 1 gram of solute per 100 milliliters final volume.

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Home

When preparing solutions from dry reagents, the same mass of any reagent is used to make a given percent concentration although the molar concentrations would

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be different. In general,  
Weight percent (w/v) = [mass  
of solute (g)/ volume of  
solution (ml)] x 100, and,  
Volume percent (v/v) =  
[volume of solute (ml)/  
volume of solution (ml)] x  
100

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~~Resource Materials: Making  
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